

alkane

saturated hydrocarbon with the general formula C_nH_{2n+2} , for example, methane, ethane, and propane

alkene

unsaturated hydrocarbon which contains a carbon–carbon double bond. Its general formula is C_nH_{2n} , for example, ethene, C_2H_4

cracking

the reaction used in the oil industry to break down large hydrocarbons into smaller, more useful ones

distillation

separation of a liquid from a mixture by evaporation followed by condensation

double bond

a covalent bond made by the sharing of two pairs of electrons

flammable

easily ignited and capable of burning rapidly

fraction

hydrocarbons with similar boiling points separated from crude oil

fractional distillation

a way to separate liquids from a mixture of liquids by boiling off the substances at different temperatures, then condensing and collecting the liquids

general formula

a formula that represents the common structure of all compounds in a single class of chemicals. For example, the general formula of all alkanes is C_nH_{2n+2}

hydrocarbon

a compound containing only hydrogen and carbon

mixture

when some elements or compounds are mixed together and intermingle but do not react together (i.e. no new substance is made). A mixture is not a pure substance

oxidised

a substance that has had oxygen added to it/ or has lost electrons

saturated hydrocarbon

describes a hydrocarbon with only single bonds between its carbon atoms. This means that it contains as many hydrogen atoms as possible in each molecule

thermal decomposition

the breakdown of a compound by heating it

unsaturated hydrocarbon

a hydrocarbon whose molecules contains at least one carbon– carbon double bond

viscosity

the resistance of a liquid to flowing or pouring; a liquid's 'thickness'